

The purpose of this summer assignment is to reacquaint you with some of the topics from IB Chemistry I that are core to the course. We will not spend a lot of time on these topics. **Completion of this assignment is optional, however there will be a test on the material during the first month of school so completion is strongly recommended.** Please print a copy of this assignment and fill in all the information in a neat and legible manner. If you have to do rough work, do it on a separate piece of paper and then transfer your work and answers to this packet. The estimated time to complete this packet approximately three hours. Resources for help are available on the last page of this assignment. Also, a key will be posted to blackboard during the first week of school to check your work. If you have questions, please feel free to email me:

Mr. Leitch: [jwleitch@fpcs.edu](mailto:jwleitch@fpcs.edu)

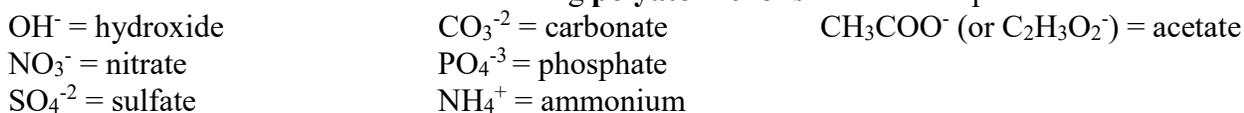
**I. Chemical Nomenclature: Remember that when we name compounds, we first must determine whether the compound is ionic or covalent.**

**Ionic compounds** follow the general formula: *Metal name nonmetal-ide*

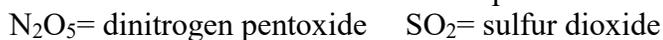
If the metal is a transition metal or any other metal under the stair case (Exceptions- Al, Zn, Ag) , we use a Roman numeral after the metal name to indicate its charge, or oxidation state. Examples of both are below:



You also need to be able to use the following **polyatomic ions** in ionic compounds:



**Covalent compounds** cannot be named in the same way due to law of multiple proportions. For example, if I said “carbon oxide”, you wouldn’t know whether I meant CO or CO<sub>2</sub>. So, we use Greek prefixes (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, and deca) to indicate the number of each type of atom. We do not use mono for the first element. Examples:



*Practice:* Fill the table with correct answers (40 Points)

Formula	Name	Name	Formula
EX: Al <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	Aluminum phosphate	EX: Manganese(IV)oxide	MnO <sub>2</sub>
K <sub>2</sub> SO <sub>4</sub>		Aluminum nitride	
N <sub>3</sub> F <sub>7</sub>		Lithium oxide	
P <sub>2</sub> O <sub>5</sub>		Phosphorous pentabromide	
CuCl <sub>2</sub>		Dichlorine heptoxide	
SO <sub>2</sub>		Silver chloride	
BaF <sub>2</sub>		Calcium carbonate	
Mg(OH) <sub>2</sub>		Nickel (II) sulfide	
AgNO <sub>3</sub>		Tin (IV) hydroxide	

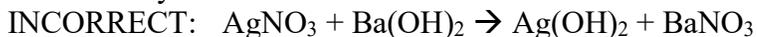
CCl <sub>4</sub>		Trisilicon tetranitride	
Mg <sub>3</sub> N <sub>2</sub>		Potassium phosphate	
NaCH <sub>3</sub> COO		Diphosphorus heptoxide	
SiCl <sub>4</sub>		Chromium (II) bromide	
N <sub>2</sub> O <sub>5</sub>		Calcium nitrate	
CoCl <sub>2</sub>		Nickel (II) oxide	
CO <sub>2</sub>		Aluminum phosphide	
CaI <sub>2</sub>		Lithium sulfide	
Be(OH) <sub>2</sub>		Nitrogen pentafluoride	
Sn(NO <sub>3</sub> ) <sub>2</sub>		Aluminum hydroxide	
P <sub>2</sub> F <sub>3</sub>		Trisilicon tetranitride	
Mg <sub>3</sub> P <sub>2</sub>		Potassium phosphate	

## II. Chemical Reactions

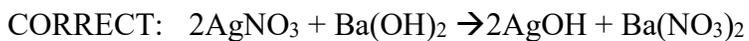
You also need to recall that there are five major types of reactions: synthesis (S), decomposition (D), single replacement (SR), double replacement (DR), and combustion (C). We will work with all of these throughout the course, and it is important that you know how to predict the products of a chemical reaction given starting materials. One of the most important things to recall is that when you write an ionic compound, the subscripts are determined by the ion charge. For example:

$\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}_2$  is INCORRECT, because it does not take into account that the charge of sodium is +1, and the charge of chloride is -1. Therefore, the correct product would be:  $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$

We balance the equation to equalize the number of atoms of each type on each side of the equation. This holds true for any ionic compound. Take the double replacement reaction between silver nitrate (AgNO<sub>3</sub>) and barium hydroxide:



This is incorrect because it does not take into account the ion charges.



*\*\*Note: later in the course we will concern ourselves with state symbols and solubility rules; this is not needed yet.* Recall that all **combustion reactions** involve a hydrocarbon reacting with oxygen to produce carbon dioxide and water vapor. (Tip: When you balance a combustion reaction please balance oxygen at the end)

*Practice:* For each of the following reactions, identify the type, predict the products and balance the equation. Assume all reactions occur. **You DO NOT need to balance problems 16-20. (40 pts)**

S.No.	Reactants	Products	Type
Ex:	$\text{Al}_2(\text{SO}_4)_3 + 3\text{Ba} \rightarrow$	$3\text{BaSO}_4 + 2\text{Al}$	SR
1.	$\text{Ag}_2\text{O} \rightarrow$		
2.	$\text{C}_4\text{H}_8 + \text{O}_2 \rightarrow$		
3.	$\text{Al} + \text{N}_2 \rightarrow$		
4.	$\text{Cu} + \text{ZnSO}_4 \rightarrow$		
5.	$\text{Pb}(\text{NO}_3)_2 + \text{K}_2\text{SO}_4 \rightarrow$		
6.	$\text{Li} + \text{O}_2 \rightarrow$		
7.	$\text{Al}_2(\text{SO}_4)_3 + \text{Ba}(\text{OH})_2 \rightarrow$		
8.	$\text{Ca} + \text{CuCO}_3 \rightarrow$		
9.	$\text{C}_3\text{H}_6 + \text{O}_2 \rightarrow$		
10.	$\text{Na}_3\text{PO}_4 + \text{Pb}(\text{NO}_3)_2 \rightarrow$		
11.	$\text{Ag}_2\text{SO}_4 + \text{AlCl}_3 \rightarrow$		
12.	$\text{CdBr}_2 + \text{Na}_2\text{S} \rightarrow$		
13.	$\text{Pb}(\text{NO}_3)_2 + \text{NaI} \rightarrow$		
14.	$\text{NaOH} + \text{Fe}(\text{NO}_3)_3 \rightarrow$		
15.	$\text{NaNO}_3 + \text{BaCl}_2 \rightarrow$		
16.	Sulfuric acid reacts with sodium hydroxide		
17.	Iron(II)chloride reacts with magnesium		
18.	Potassium bromide reacts with calcium oxide		
19.	Rubidium reacts with sulfur		
20.	Diphosphorus pentaoxide decomposes		

### III. Stoichiometry (40 pts)

I expect that you will enter this class with a grasp of the fundamentals of stoichiometry. We will spend one class period reviewing gas-specific stoichiometry (including the concept of molar volume) and one class working on problems with limiting reactants. If you need extra help recalling some of the other material, I will be happy to help! I also expect that you will use DIMENSIONAL ANALYSIS (aka “the fence”) in my class. If you did not learn this method, I will work with you to solve problems this way. The steps to using dimensional analysis as well as some worked examples are given below.

**Mole conversions:** Complete each of the following problems. CIRCLE YOUR FINAL ANSWER!! Be sure to include units and show all work!! (4pts each)

EX: Find the mass (in grams) of  $3.65 \times 10^{-2}$  mol potassium carbonate

$$\frac{3.65 \times 10^{-2} \text{ mol K}_2\text{CO}_3}{1} \times \frac{126 \text{ g K}_2\text{CO}_3}{1 \text{ mole K}_2\text{CO}_3} = 4.60 \text{ g K}_2\text{CO}_3$$

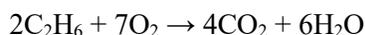
EX: How many particles of chlorine molecules are present in 50.0g of Chlorine?

$$\frac{50.0 \text{ grams Cl}_2}{71.0 \text{ grams Cl}_2} \times \frac{1 \text{ mole Cl}_2}{1} \times \frac{6.02 \times 10^{23} \text{ molecules SO}_2}{1 \text{ mole Cl}_2} = 4.24 \times 10^{23} \text{ molecules Cl}_2$$

- Convert  $9.3 \times 10^{15}$  atoms of lead to moles of lead.
- Mass of  $1.906 \times 10^{-2}$  mol  $\text{BaI}_2$
- Number of O atoms in  $4.88 \times 10^{-3}$  mol  $\text{Al}(\text{NO}_3)_3$
- Find the number of atoms in 45.6 g of sulfur.

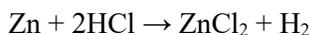
**Stoichiometry:** Complete each of the following problems. CIRCLE YOUR FINAL ANSWER!! Be sure to include units and show all work!! (4 pts each)

EX: To produce 12 moles of water, how many moles of oxygen gas are needed?



$$\frac{12 \text{ mole H}_2\text{O}}{6 \text{ mole H}_2\text{O}} \times \frac{7 \text{ mole O}_2}{1} = 14 \text{ mole O}_2$$

EX: If 0.600 gram of zinc is used, how many grams of zinc chloride will be produced in the above reaction?



$$\frac{0.0600 \text{ grams Zn}}{65 \text{ grams Zn}} \times \frac{1 \text{ mole Zn}}{1} \times \frac{1 \text{ mole ZnCl}_2}{1} \times \frac{136 \text{ grams ZnCl}_2}{1 \text{ mole ZnCl}_2} = 1.255 \text{ grams ZnCl}_2$$

- Carbon dioxide is released into the atmosphere through the combustion of octane ( $\text{C}_8\text{H}_{18}$ ) in gasoline. Calculate the mass of octane needed to release 5.00 mol  $\text{CO}_2$ .



6. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.

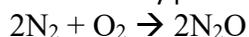


7. Calculate the mass of chlorine gas needed to react with 84.0 grams of aluminum to produce aluminum chloride. **Write a balanced chemical equation first!**

8. An iron ore sample contains iron (III) oxide. It reacts with CO to produce iron metal. Calculate the number of grams of CO needed to react with 0.150 kg of iron (III) oxide. *\*Note: balance the equation first!!*  $\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$

#### Limiting and excess reactant (8 pts)

9. What mass of "laughing gas" (dinitrogen monoxide) will be produced from 50 g of nitrogen gas and 75 g of oxygen gas? (Hint: Two stoichiometry problems in one question)



#### Helpful Resources:

1. Mrs. Wiseman's IB Chem (great site with video tutorials and extra practice!):  
[http://www.mwiseman.com/courses/chem\\_ib/](http://www.mwiseman.com/courses/chem_ib/)
2. Chem I textbook online:  
[www.chemistrymc.com](http://www.chemistrymc.com), then click online student textbook full version and enter password BAC17E706D
3. Helpful videos:

#### **Naming ionic compounds**

<https://www.youtube.com/watch?v=mKo72RnN37E>

#### **Naming Molecular (Covalently Bonded) Compounds**

<https://www.youtube.com/watch?v=DejkrR4pvRw>

#### **Stoichiometry:**

<https://www.youtube.com/watch?v=SjQG3rKSZUQ>

4. Your chem. I notes...you still have them...right?? ☺